C1–Li2′ is 51.8 (3)° bent out of the plane of the cyclopropane ring (Li2′ lies 168.5 (8) pm below this plane). This bonding situation differs characteristically from that of the α-cyanobenzyllithium species **4**: In **4** the anionic α-C atom is planar and has no carbon–lithium bond, and the C–C–N unit is linear. The tetrahedral configuration of C1 in [**3**·THF]$_\infty$ together with the higher acidity of cyclopropyl as compared to isopropyl nitrile ("cyclopropyl effect")[2,4] thus clearly demonstrates that the (inductive) field effect plays a major role in the stabilization of a negative charge by a nitrile group.[6b] In agreement the bond lengths C1–C6 (140.0 (0.7) pm) and C6–N1 (117.8 (0.7) pm) are not very different from those in cyclopropyl nitriles (e.g., 1,1,2,2-tetracyanocyclopropane C–C 144.2 pm; C–N 115.0 pm).[7] In a keteniminate type structure resulting from delocalization of the negative charge one would expect a shorter C–C and a longer C–N bond as this is the case in the ketenimine *N*-*p*-bromophenyl-2,2-diphenylvinylideneamine: C–C 133.0 pm; C–N 124.0 pm.[8] The bond lengths within the three-membered ring are also of interest: The distal C2–C3 bond (147.3 (0.7) pm) is shorter than the two vicinal bonds (C1–C2 150.0 (0.7) pm; C1–C3 152.2 (0.7) pm). Such a pattern of bond length asymmetry has been predicted for acceptor-substituted *cyclopropanes* if there is an interaction between the occupied cyclopropane 3e′ and an unoccupied acceptor π orbital.[9,10] It is generally observed in cyano-substituted cyclopropanes.[7] Since this is also observed in the cyanocyclopropyl anion [**3**·THF]$_\infty$, the nitrile group not only interacts with the "anionic" but also with the 3e′ orbital.[11]

The tetrahedral configuration of the anionic C1 in the *solid-state structure* of [**3**·THF]$_\infty$ thus nicely confirms the results of H/D exchange experiments with cyclopropylnitriles in *solution*.[2,4] It is also in agreement with *calculations* which predicted a tetrahedral anionic carbon atom in the α-cyanocyclopropyl anion.[12] Furthermore, MNDO calculations indicated that the most stable unsolvated dimer of LiCH$_2$CN prefers an eight-membered ring structure, while solvation favors the alternative Li–N–Li–N four-membered ring.[13] Because of the special features of an α-cyanocyclopropyl anion as outlined above, both of these structural units are verified in [**3**·THF]$_\infty$.

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Supplementary Material Available: Tables of atomic positional and isotropic equivalent thermal parameters, anisotropic thermal parameters, bond distances, bond angles, torsion angles, interatomic contacts greater than 200 and less than 300 pm, and weighted least-squares planes and lines for [**3**·THF]$_\infty$ (10 pages). Ordering information is given on any current masthead page.

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Additions and Corrections

Acromelic Acids A and B. Potent Neuroexcitatory Amino Acids Isolated from *Clitocybe acromelalga* [*J. Am. Chem. Soc.* **1988**, *110*, 4807–4815]. KATSUHIRO KONNO, KIMIKO HASHIMOTO, YASUFUMI OHFUME, HARUHISA SHIRAHAMA,* and TAKESHI MATSUMOTO

Page 4807: The following should be added to ref 3. Syntheses of acromelic acid A were reported from other laboratories: Takano, S.; Iwabuchi, Y.; Ogasawara, K. *J. Am. Chem. Soc.* **1987**, *109*, 5523–5524. Baldwin, J. E.; Li, Chun-Sing *J. Chem. Soc., Chem. Commun.* **1988**, 261–263.

Effect of Long-Distance Electron Transfer on Chemiluminescence Efficiencies [*J. Am. Chem. Soc.* **1988**, *110*, 2764–2772]. ROBERT D. MUSSELL and DANIEL G. NOCERA*

Page 2767: The $E_{1/2}$ values for the reduction potential of acceptors (A) listed in Table I are given vs the ferrocenium/ferrocene reference couple and not as written, vs SCE. To relate these potentials vs a SCE reference, a value of 0.31 V must be added to the listed potentials.

Page 2769: Equation 14 should read $H_{AB}^2 = (H_{AB}^0)^2 \exp[-\beta(r - \delta)]$.

Synthesis and Reactivity of Binuclear Tropocoronand and Related Organocopper(I) Complexes. Catalytic Enantioselective Conjugate Addition of Grignard Reagents to 2-Cyclohexen-1-one [*J. Am. Chem. Soc.* **1988**, *110*, 3175]. GILBERTO M. VILLACORTA, CH. PULLA RAO, and STEPHEN J. LIPPARD*

In writing this paper we were unaware of the previous synthesis of an optical isomer of the ligand that we call H(CHIRAMT) (Brunner, H.; Knott, A.; Benn, R.; Rufinska, A. *J. Organomet. Chem.* **1985**, *295*, 211). We thank Dr. H. Brunner for drawing this prior work to our attention.

A Carbon–Carbon Bond Cleavage Reaction of Carbon Suboxide at a Metal Center. Synthesis and Structural Characterization of $\mathbf{WCl_2(CO)(PMePh_2)_2\{C,C'{:}\eta^2\text{-}C(O)CPMePh_2\}}$ [*J. Am. Chem. Soc.* **1988**, *110*, 4855]. ADAM K. LIST, GREGORY L. HILLHOUSE,* and ARNOLD L. RHEINGOLD*

Compound **3** is properly formulated as the phosphinocarbyne complex $WCl_2(CO)(CPMePh_2)(PMePh_2)_2$, and not $WCl_2(CO)(PMePh_2)_3$ as was reported. This was confirmed by an X-ray structural determination on a PPh_3 derivative, $WCl_2(CO)(CPPh_3)(PMePh_2)_2$, details of which will be published elsewhere.